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We have amount methane and need molecules of hydrogen, so we use molar ratio of moles hydrogen to moles methane. Get the molar mass of methane by adding $1 \times$ molar mass carbon + $4 \times$ molar mass of hydrogen. Using the initial amount, divide it by number of grams in 1 mole. Then, multiply that by the molar

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11 SOLUTIONS MANUAL Section 11.1 Defining
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371 – 372 1. Interpret the following balanced chemical
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that the law of conservation of mass is observed. a. N_2 (g)
 3H_2 (g) 2NH_3 (g) 1 ...

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